



4 mole of a mixture of molar salt and  $\text{Fe}_2\text{SO}_4$  requires 500 ml of  $\text{KMnO}_7$  for complete oxidation in acidic medium the mole of the molar salt in the mixture is

in a process temperature and volume of one mole of an ideal monoatomic gas are varied according to the relation  $v \propto T^k$  where  $k$  is a constant in this process the temperature of the gas is increased

mole can be defined as the amount of substance that contains the same number of entities atoms molecules ions etc as the number of atoms present in 12g of the carbon 12 isotope

one mole of an ideal gas at temperature  $T_1$  expands according to the law  $p \propto V^{-2}$  a constant the work done by the gas till the temperature of gas become  $T_2$  is

the Henry's law constant for the solubility of  $\text{N}_2$  gas in water at 298 K is  $1.0 \times 10^5$  atm the mole fraction of  $\text{N}_2$  in air is 0.8 the number of moles of  $\text{N}_2$  from air dissolved in 10 moles of water at 298 K and

an ideal solution is formed by mixing two volatile liquids A and B  $x_A$  and  $x_B$  are the mole fractions of A and B respectively in the solution and  $y_A$  and  $y_B$  are the mole fractions of A and B respectively in

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